

Name \_\_\_\_\_

Hour \_\_\_\_\_

## Redox - Balancing Half Reactions

\* For each of the following half reactions, identify each as oxidation or reduction and balance them to make sure that the correct number of electrons

	Reaction Type	Balanced half reaction
1. $\text{Cu}_{(s)} \rightarrow \text{Cu}^{2+}_{(aq)}$		
2. $3\text{O}_2(g) \rightarrow 6\text{O}^{2-}_{(aq)}$		
3. $\text{Al}_{(s)} \rightarrow \text{Al}^{3+}_{(aq)}$		
4. $\text{Fe}^{2+}_{(aq)} \rightarrow \text{Fe}_{(s)}$		
5. $\text{Cd}^{2+}_{(aq)} \rightarrow \text{Cd}_{(s)}$		
6. $2\text{Ag}^{+}_{(aq)} \rightarrow 2\text{Ag}_{(s)}$		
7. $2\text{Fe}^{3+}_{(aq)} \rightarrow 3\text{Fe}^{2+}_{(aq)}$		
8. $3\text{Cl}_2(g) \rightarrow 6\text{Cl}^{-}_{(aq)}$		
9. $8\text{I}^{-}_{(aq)} \rightarrow 4\text{I}_{2(s)}$		
10. $3\text{Cr}^{2+}_{(aq)} \rightarrow 3\text{Cr}^{3+}_{(aq)}$		
11. $2\text{Br}^{-}_{(aq)} \rightarrow \text{Br}_{2(g)}$		
12. $4\text{Cu}^{2+}_{(aq)} \rightarrow 4\text{Cu}_{(s)}$		
13. $2\text{Pb}_{(s)} \rightarrow 2\text{Pb}^{2+}_{(aq)}$		
14. $\text{Sc}^{3+}_{(aq)} \rightarrow \text{Sc}_{(s)}$		
15. $5\text{Au}_{(s)} \rightarrow 5\text{Au}^{+}_{(aq)}$		
16. $4\text{Cd}^{2+}_{(aq)} \rightarrow 4\text{Cd}_{(s)}$		
17. $3\text{Mn}^{2+}_{(aq)} \rightarrow 3\text{Mn}_{(s)}$		