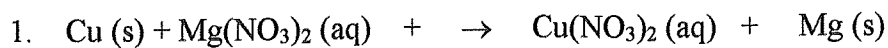


Redox Reactions

\* Write the T.I.E.'s and N.I.E.'s for each of the following molecular equations – and then determine the balanced half reactions for the oxidation and reduction that are taking place.

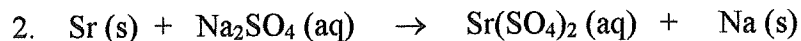


T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction

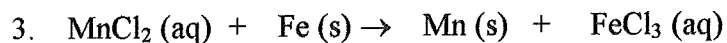


T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction

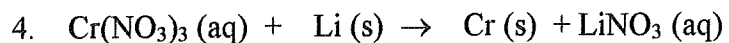


T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction

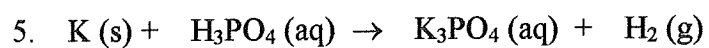


T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction

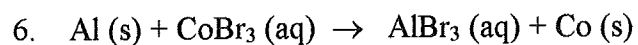


T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction



T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction