Name	
	Date

## **Redox Reactions**

\* Write the <u>T.I.E.'s</u> and <u>N.I.E.'s</u> for each of the following molecular equations – and then determine the <u>balanced half reactions</u> for the oxidation and reduction that are taking place.

1.  $Cu(s) + Mg(NO_3)_2(aq) + \rightarrow Cu(NO_3)_2(aq) + Mg(s)$ 

T.I.E. =

N.I.E. =

**Balanced Oxidation Half-reaction** 

Balanced Reduction Half-reaction

2.  $Sr(s) + Na_2SO_4(aq) \rightarrow Sr(SO_4)_2(aq) + Na(s)$ 

T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction

3.  $MnCl_2(aq) + Fe(s) \rightarrow Mn(s) + FeCl_3(aq)$ 

T.I.E. =

N.I.E. =

Balanced Oxidation Half-reaction

Balanced Reduction Half-reaction